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# International Standard



# 6994

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INTERNATIONAL ORGANIZATION FOR STANDARDIZATION • МЕЖДУНАРОДНАЯ ОРГАНИЗАЦИЯ ПО СТАНДАРТИЗАЦИИ • ORGANISATION INTERNATIONALE DE NORMALISATION

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## Aluminium ores — Determination of aluminium content — EDTA titrimetric method

*Minerais alumineux — Dosage de l'aluminium — Méthode titrimétrique à l'EDTA*

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## Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work.

Draft International Standards adopted by the technical committees are circulated to the member bodies for approval before their acceptance as International Standards by the ISO Council. They are approved in accordance with ISO procedures requiring at least 75 % approval by the member bodies voting.

International Standard ISO 6994 was prepared by Technical Committee ISO/TC 129, *Aluminium ores*.

Users should note that all International Standards undergo revision from time to time and that any reference made herein to any other International Standard implies its latest edition, unless otherwise stated.

# Aluminium ores — Determination of aluminium content — EDTA titrimetric method

## 1 Scope and field of application

This International Standard specifies an EDTA titrimetric method for the determination of the aluminium content of aluminium ores.

The method is applicable to products having aluminium contents, expressed as aluminium oxide ( $\text{Al}_2\text{O}_3$ ), in the range 30 to 65 % (*m/m*).

## 2 Reference

ISO 8557, *Aluminium ores — Determination of hygroscopic moisture in analytical samples — Gravimetric method.*

## 3 Principle

Decomposition of the test portion by either

- a) treatment with hydrochloric, nitric and sulfuric acids;

NOTE — This method is recommended for ores containing gibbsite and/or boehmite and when the residue from the dissolution of the test portion after silica volatilization is less than 1 % of the mass of the test portion.

or

- b) sintering with sodium peroxide followed by brief fusion and dissolution of the melt with sulfuric acid;

NOTE — This method is recommended for ores containing diaspore and when the residue from the dissolution of the test portion after silica volatilization is greater than 1 % of the test portion.

or

- c) fusion with sodium tetraborate/sodium carbonate flux followed by dissolution of the melt with sulfuric acid.

NOTE — This method is suitable for all types of ore.

Removal of silica by filtration to give a solution containing aluminium and other elements. Measurement of an aliquot portion and precipitation of aluminium and other elements except chromium, manganese and vanadium, by ammonium benzoate in acid solution. Filtration and dissolution of the precipitated benzoates in acid followed by precipitation of elements except aluminium by sodium hydroxide. Filtration of the precipitate followed by acid dissolution, reprecipitation with sodium hydroxide and filtration to recover coprecipitated aluminium. Addition of an excess of EDTA to the combined filtrates and determination of aluminium content by back titration of the excess EDTA with zinc sulfate solution using xylenol orange as indicator.

## 4 Reagents

During the analysis, use only reagents of analytical reagent grade and only distilled water or water of equivalent purity.

### 4.1 Sodium peroxide ( $\text{Na}_2\text{O}_2$ ), powdered.

NOTE — Sodium peroxide should be protected against humidity and should not be used once it has begun to agglomerate.

### 4.2 Sodium carbonate/sodium tetraborate flux.

Mix intimately 3 parts by mass of anhydrous sodium carbonate ( $\text{Na}_2\text{CO}_3$ ) and 1 part by mass of anhydrous sodium tetraborate ( $\text{Na}_2\text{B}_4\text{O}_7$ ).

### 4.3 Ammonium acetate, 250 g/l solution.

Dissolve 250 g of ammonium acetate in 900 ml of water and adjust the pH to 6 by the addition of glacial acetic acid ( $\rho$  1,050 g/ml). Dilute to 1 litre.

### 4.4 Ammonium benzoate, 100 g/l solution.

NOTE — To obtain a clear solution after warming, it may be necessary to add up to 5 ml of ammonia solution (4.6). Filter if necessary to remove any particulate residue.

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### 4.5 Hexamethylenetetramine (hexamine), 250 g/l solution.

Dissolve 25 g of hexamethylenetetramine in 70 ml of water and dilute to 100 ml.

### 4.6 Ammonia solution, ( $\rho$ 0,9 g/ml), diluted 1+1.

### 4.7 Glacial acetic acid, ( $\rho$ 1,050 g/ml).

### 4.8 Hydrochloric acid ( $\rho$ 1,17 g/ml), diluted 2+3.

### 4.9 Sodium hydroxide, 160 g/l solution.

This solution shall be stored in a plastic bottle.

### 4.10 Sulfuric acid ( $\rho$ 1,84 g/ml), diluted 1+1.

### 4.11 Sulfuric acid ( $\rho$ 1,84 g/ml), diluted 1+3.

### 4.12 Sulfuric acid ( $\rho$ 1,84 g/ml), diluted 1+9.

### 4.13 Hydrofluoric acid, 40 % (m/m), ( $\rho$ 1,13 g/ml).

### 4.14 Sodium sulfate, ( $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ ), 100 g/l solution.

### 4.15 Potassium permanganate, 1 g/l solution.

### 4.16 Acid mixture.

Measure into a 1 litre beaker 225 ml of water and add cautiously, with stirring, 175 ml of sulfuric acid ( $\rho$  1,84 g/ml). Cool to room temperature and add 150 ml of hydrochloric acid ( $\rho$  1,17 g/ml) and 50 ml of nitric acid ( $\rho$  1,42 g/ml).

This solution shall be prepared for each series of determinations and any unused solution discarded.

### 4.17 Standard aluminium solution.

Dissolve approximately 1,00 g, weighed to the nearest 0,000 1 g, of aluminium metal (not less than 99,9 %) in 60 ml of the hydrochloric acid (4.8), heating if necessary. Cool and dilute to the mark in a 1 000 ml one-mark volumetric flask.

Calculate the equivalent aluminium oxide concentration  $a$  of the standard solution, in grams per litre, using the equation

$$a = m_1 \times 1,8895 \times \frac{P}{100}$$

where

$m_1$  is the mass of aluminium metal, in grams;

$P$  is the purity of aluminium used, in per cent.

### 4.18 Ethylenediaminetetraacetic acid disodium salt dihydrate (EDTA), approximately 0,05 mol/l (18,6 g/l) standard volumetric solution.

### 4.19 Zinc sulfate, standard volumetric solution, $c(\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}) = 0,05$ mol/l.

Dissolve 14,5 g of zinc sulfate heptahydrate ( $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$ ) in water, transfer to a 1 000 ml one-mark volumetric flask, dilute to volume and mix.

#### 4.19.1 Standardization procedure

##### 4.19.1.1 Titration of EDTA solution with standard zinc sulfate solution

Using a pipette, add 50 ml of the EDTA solution (4.18) to a 500 ml beaker and add 250 ml of water. Add 0,3 ml of the methyl orange indicator (4.20) and, if necessary, add the ammonia solution (4.6) drop by drop until the colour just changes to yellow.

Add 20 ml of the ammonium acetate solution (4.3) and adjust to  $\text{pH } 5,7 \pm 0,1$ , using either the ammonia solution (4.6) or the glacial acetic acid (4.7). Add 0,5 ml of the xylenol orange indicator (4.21) and titrate to the end-point (colour change yellow to purple) with zinc sulfate solution (4.19). Record the titre ( $V_1$ ).

##### 4.19.1.2 Determination of the equivalence factor for zinc sulfate solution

Pipette 50 ml ( $V_3$ ) of the standard aluminium solution (4.17) into a 500 ml beaker, and add 200 ml of water.

With a pipette, add 50 ml of the EDTA solution (4.18), and 0,3 ml of the methyl orange indicator (4.20). Add the ammonia solution (4.6) drop by drop until the colour just changes.

Add 20 ml of the ammonium acetate solution (4.3) and a boiling aid. Boil for 10 min and cool. Adjust to  $\text{pH } 5,7 \pm 0,1$ , as in 4.19.1.1. Add 0,5 ml of the xylenol orange indicator (4.21) and titrate to the end-point with the zinc sulfate solution (4.19). Record the titre ( $V_2$ ).

Calculate the equivalence factor  $A$  of the zinc sulfate solution, using the formula

$$1 \text{ ml zinc sulfate solution } (A) \equiv \frac{V_3 a}{V_1 - V_2} \text{ mg of Al}_2\text{O}_3$$

where  $a$  is the mass, in grams, of aluminium oxide per litre of standard aluminium solution (4.17).

### 4.20 Methyl orange indicator, 0,1 g/l solution.

### 4.21 Xylenol orange indicator, 1 g/l solution.

Prepare each month and store in a dark bottle.

## 5 Apparatus

Ordinary laboratory apparatus, and

**5.1 Dishes**, of inert material and sufficient size to take the required quantity of sample at a layer density of 5 mg/mm<sup>2</sup>.

**5.2 Oven**, electrically heated with a temperature-control device and mechanical convection, capable of maintaining a temperature of 105 ± 5 °C.

NOTE — The oven should be checked by placing the usual number of drying dishes in position in the oven and then measuring the temperature in the vicinity of each dish to ensure that the required temperature is maintained throughout the oven, not just in the neighbourhood of the controlling element.

**5.3 Platinum/gold (95/5) crucibles**, of approximately 25 ml capacity for the fusion decomposition in 7.6.1.3.

**5.4 Zirconium crucibles**, of approximately 40 ml capacity for the sinter decomposition in 7.6.1.2.

**5.5 Platinum crucibles.**

**5.6 Filter papers.**

**5.6.1 Medium-textured filter papers.**

NOTE — Whatman 540 grade have been found suitable.

**5.6.2 Fast, hardened, ashless filter papers.**

NOTE — Whatman 541 grade have been found suitable.

## 6 Sampling and samples

### 6.1 Sampling

Laboratory samples shall be taken and ground to pass a 150 µm test sieve by the procedures specified in the relevant standards.<sup>1)</sup>

### 6.2 Preparation of the test sample

Take approximately 10 g of the laboratory sample and place in the sample dish (5.1). Spread the sample evenly to give a layer density of approximately 5 mg/mm<sup>2</sup> and allow the sample to equilibrate with the laboratory atmosphere for a minimum of 2 h.

## 7 Procedure

### 7.1 Number of determinations

Carry out the analysis in duplicate and independently on one ore sample.

NOTE — The expression "independently" implies a change of the person carrying out the analysis. If the same person must carry out the analysis, the procedure should be carried out at different times.

### 7.2 Test portion

Weigh, to the nearest 0,000 1 g, approximately 1 g of air-equilibrated sample and record the mass (*m*).

NOTE — For ores containing more than 55 % (*m/m*) of aluminium oxide, the mass of the test portion should be 0,9 g.

### 7.3 Determination of hygroscopic moisture content

Simultaneously with the weighing of the test portion, weigh two test portions for the determination of hygroscopic moisture and carry out the determinations by the procedure specified in ISO 8557.

### 7.4 Blank test

In parallel with the decomposition of the test portion, prepare a blank test solution appropriate to the method of decomposition but with the omission of the test portion. When the analysis is carried out on several samples at the same time, the blank test value may be represented by one test, provided that the procedure is the same and the reagents used are from the same bottles.

### 7.5 Check test

In each run, one analysis of a certified reference material of the same type of ore shall be carried out in parallel and under the same conditions with the analysis of one ore sample.

NOTE — The certified reference material should be of the same type as the sample to be analysed. Such certified reference materials cannot, however, be considered as being of the same type if the properties of the sample to be analysed differ from those of the certified reference material to such an extent that the analytical procedure must be changed substantially.

### 7.6 Preparation of the test solution

#### 7.6.1 Decomposition of the test portion

If the decomposition is to be based on acid attack, proceed as specified in 7.6.1.1. If the decomposition is to be based on alkali sintering, proceed as specified in 7.6.1.2. If the decomposition is to be based on fusion, proceed as specified in 7.6.1.3.

##### 7.6.1.1 Acid decomposition

Transfer the test portion (7.2) to a 400 ml beaker. Moisten with water and add 60 ml of acid mixture (4.16). Cover the beaker and heat at approximately 80 °C to decompose the sample (see note 1).

1) Where no International Standards exist, the relevant standards shall be the national standards. Two International Standards on this subject are currently in preparation: ISO 6137, *Aluminium ores — Method of sampling*, and ISO 6140, *Aluminium ores — Preparation of samples*.

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When the evolution of brown fumes ceases, carefully rinse the cover and wall of the beaker. Evaporate the uncovered solution to dense fumes of sulfuric acid. Replace the cover and heat strongly for 60 min on a hotplate set to produce in the solution a temperature of  $210 \pm 10$  °C (see note 2).

### NOTES

1 With samples of high iron content [ $> 15$  % (*m/m*) of  $\text{Fe}_2\text{O}_3$ ], the initial heating should be conducted for an extended period.

2 Determine the temperature setting by using a separate test beaker containing a partial immersion thermometer in 10 mm depth of sulfuric acid ( $\rho$  1,84 g/ml).

### 7.6.1.2 Alkali sinter decomposition

Transfer the test portion (7.2) to a dry zirconium crucible (5.4), add 10 g of the sodium peroxide (4.1) and mix intimately with a dry metal spatula. Place in a muffle furnace at 480 to 500 °C for 45 min. Remove the crucible and contents from the furnace and heat over a burner to melt the sinter (about 30 s). Continue heating the melt while swirling for a total time of 2 min.

Allow the crucible to cool to room temperature (a metal block may be used to expedite this process). If necessary, clean the base of the crucible by wiping with a damp filter paper. Place the crucible on its side in a 400 ml beaker, cover and add cautiously to the beaker 140 ml of the sulfuric acid (4.12). Add 20 ml of the sulfuric acid (4.10) and digest to leach the contents of the crucible. Remove the crucible when the contents are completely leached and rinse well into the solution in the beaker.

Carefully rinse the cover and wall of the beaker. Evaporate the uncovered solution to dense fumes of sulfuric acid. Replace the cover and heat strongly for 60 min on a hotplate set to produce in the solution a temperature of  $210 \pm 10$  °C.

NOTE — Determine the temperature setting by using a separate test beaker containing a partial immersion thermometer in a 10 mm depth of sulfuric acid ( $\rho$  1,84 g/ml).

### 7.6.1.3 Fusion decomposition

Transfer the test portion (7.2) to a dry platinum/gold crucible (5.3) (see note 1), add 2 g of the sodium carbonate/sodium tetraborate flux (4.2) and mix intimately with a dry metal spatula. Cover with a platinum lid and place in a furnace at 480 to 500 °C for 5 min, then transfer to a furnace at 1 000 °C for 15 min.

Allow the crucible to cool and place it on its side in a 400 ml beaker. Add 70 ml of the sulfuric acid (4.11), cover and digest to leach the contents of the crucible. Remove the crucible when the contents are completely leached and rinse well into the main solution.

Carefully rinse the cover and wall of the beaker. Evaporate the uncovered solution to dense fumes of sulfuric acid. Completely replace the cover and heat strongly for 60 min on a hotplate set to produce in the solution a temperature of  $210 \pm 10$  °C (see note 2).

### NOTES

1 A platinum crucible may be used but care should be taken to ensure that the hot fusion mixture does not 'creep' over the side of the crucible. Consequently, it would be preferable to use a larger platinum crucible (e.g. 30 ml).

2 Determine the temperature setting by using a separate test beaker containing a partial immersion thermometer in a 10 mm depth of sulfuric acid ( $\rho$  1,84 g/ml).

### 7.6.2 Dissolution and filtration

Cool the solution from 7.6.1.1, 7.6.1.2 or 7.6.1.3, to room temperature, then cautiously add 130 ml of water and heat to 80 to 90 °C for at least 40 min, with stirring, to completely dissolve the salts. Add a small quantity of ashless pulp with stirring and filter while hot through a medium-textured filter paper, (5.6.1) collecting the filtrate in a 250 ml one-mark volumetric flask. Rinse the beaker with water, and, using a rubber-tipped glass rod, clean the beaker and transfer the residue quantitatively to the paper. Wash the paper and residue five times with 5 to 10 ml aliquot portions of hot water. Reserve the filtrate and washings. Retain the filter paper and residue and continue according to 7.6.3.

### 7.6.3 Treatment of the residue

Transfer the filter paper and residue from 7.6.2 to a pre-ignited tared platinum crucible. Dry and char slowly and then ash the filter paper by increasing the temperature to 600 to 700 °C for 30 min in a muffle furnace. Allow the crucible to cool and moisten the residue with a few drops of water. Add 5 drops of the sulfuric acid solution (4.10) and 8 ml of the hydrofluoric acid (4.13). Evaporate carefully to dryness in a fume cupboard to expel silica and sulfuric acid. Heat gently over a low flame, then to red heat, allow the crucible to cool, and weigh. Determine the mass of residue to ensure that the requirements of clause 3 are met.

Add  $0,7 \pm 0,1$  g of the flux (4.2). Fuse at above 1 000 °C for 4 to 5 min in a muffle furnace, roll the melt briefly then return to the furnace for 1 to 2 min.

NOTE — If necessary for complete attack, a gas burner may be beneficial.

Allow the crucible to cool, add 10 ml of the sulfuric acid solution (4.12), cover and warm to dissolve the fused mass. Transfer the solution to the reserved filtrate (7.6.2), rinsing the crucible with water. Cool to room temperature, dilute to the mark with water and mix. This is the test solution.

## 7.7 Determination of aluminium content

Pipette a 50,0 ml aliquot portion of the test solution into a 500 ml beaker. Dilute to 120 ml. Heat the aliquot portion to 90 °C and add the potassium permanganate solution (4.15) drop by drop until the purple colour persists for at least 10 s. Continue heating at 90 °C for 5 min.

Cool to room temperature. Using a pH-meter, add the ammonia solution (4.6), with stirring, to give a pH of  $2,4 \pm 0,1$ . Add slowly 15 ml of the ammonium benzoate solution (4.4), and 4 ml of the hexamine solution (4.5), then heat at 90 °C for 30 min, stirring every 10 min. Do not boil, as this will cause frothing. Cool to about 60 °C, filter through an 11 cm fast, hardened, ashless filter paper (5.6.2) and reject the filtrate.

NOTE — Avoid allowing precipitate to dry on the beaker wall and lip by immediately washing down with water any residual precipitate not transferred to the filter. Clean the wall with a rubber-tipped glass rod and retain any precipitate removed in the beaker. Washing of the precipitate on the filter is not required.

Add 3 ml of the hydrochloric acid (4.8) to the original beaker, cover the beaker and reflux the acid to dissolve residues of precipitate on the wall. Using a rubber-tipped glass rod, ensure that all residual aluminium is dissolved from the beaker wall and the lip area. Rinse and remove the rod. Transfer all the precipitate in the funnel to the original beaker with water, ensuring that the paper remains in the funnel, and place the beaker beneath the funnel. Using a 2 ml pipette, thoroughly wash the paper and funnel with a maximum of 12 ml of the hot hydrochloric acid (4.8), followed by hot water. Discard the paper and warm the solution to 60 °C to dissolve the precipitate. Add, with stirring, 25 ml of the sodium hydroxide solution (4.9) and continue heating at 50 to 60 °C for 10 min, stirring occasionally. Filter into a 500 ml beaker containing 6 ml of the sulfuric acid (4.10) through an 11 cm fast, hardened filter paper (5.6.2). Add a stirring rod to the filtrate and mix the acid and the filtrate from time to time. Wash the precipitate into the funnel with sodium sulfate solution (4.14) using a rubber-tipped glass rod to clean the beaker wall and wash the precipitate and paper three times with the sodium sulfate wash solution. Reserve the filtrate containing the aluminium. This is the reserved filtrate.

Add 3 ml of the hydrochloric acid (4.8) to the original beaker, cover the beaker and reflux the acid to dissolve residues of precipitate on the wall. Transfer the precipitate in the funnel to the original beaker with water and place the beaker under the funnel containing the paper. Dissolve the remaining precipitate on the paper by careful washing of the whole of the paper with 12 ml of the hot hydrochloric acid (4.8), followed by hot water. Discard the paper and warm the solution to dissolve the precipitate. Heat to 50 to 60 °C, add, with stirring, 25 ml of the sodium hydroxide solution (4.9) and continue heating at 50 to 60 °C for 10 min.

Filter into the beaker containing the reserved filtrate, through an 11 cm fast, hardened filter paper (5.6.2). Transfer the precipitate to the paper and wash the paper and precipitate several times with sodium sulfate wash solution (4.14).

Heat the combined filtrate to 60 to 70 °C and add by pipette 50,0 ml of the EDTA solution (4.18). Add 0,3 ml of the methyl orange indicator (4.20) and neutralize just to the orange colour

with the ammonia solution (4.6). Add 20 ml of the ammonium acetate solution (4.3), boil for 10 min and cool. Adjust to pH  $5,7 \pm 0,1$  with the ammonia solution (4.6) or glacial acetic acid (4.7). Add 0,5 ml of the xylenol orange indicator solution (4.21) and titrate with the zinc sulfate solution (4.19). Record the titres  $V$  and  $V_b$  for the test and blank test solutions respectively.

## 8 Expression of results

### 8.1 Calculation of aluminium content

Calculate the aluminium content, expressed as a percentage by mass as aluminium oxide, using the formula

$$\frac{250 A (V_b - V)}{10 V_p m} \times K$$

where

$A$  is the equivalence factor for zinc sulfate solution derived in 4.19.1.2;

$V_b$  is the volume, in millilitres, of zinc sulfate solution (4.19) used to titrate the blank test solution;

$V$  is the volume, in millilitres, of zinc sulfate solution (4.19) used to titrate the test solution;

$V_p$  is the volume, in millilitres, of the aliquot portion taken in 7.7;

$m$  is the mass, in grams, of the test portion;

$K$  is the conversion factor found from the formula

$$K = \frac{100}{100 - H}$$

where  $H$  is the hygroscopic moisture content, as a percentage by mass, determined in accordance with ISO 8557.

### 8.2 General treatment of results

#### 8.2.1 Precision

A planned trial was carried out by ten countries with a total of sixteen laboratories participating. Two determinations were carried out on each of four samples. The repeatability, reproducibility and reproducibility index have been calculated and are presented in the table.

#### 8.2.2 Acceptance of analytical values

The analytical value for the test sample shall be accepted when the analytical value obtained for the corresponding certified reference material agrees with the certified value within the limit of the reproducibility index contained in the table, and when the difference between the two values for the test sample does not exceed  $2,77 s_w$  as calculated from the appropriate value of  $s_w$  given in the table.

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When the analytical value obtained for the certified reference material is outside the reproducibility index, an analysis shall be carried out simultaneously on one test sample and one certified reference material, with one blank test. The analytical value obtained for the certified reference material shall be examined for the acceptance of the value for the test sample, as above. If the value obtained for the certified reference material is again outside the limits, the procedure shall be repeated with a different reference material of the same type of ore until two acceptable values for the test sample are obtained.

When the range (absolute difference) of the two values for the test sample is greater than  $2,77 s_w$ , additional analyses shall be carried out on one test sample with one blank test simultaneously with an analysis of a certified reference material of the same type of ore. Acceptance of such additional values for the test sample shall be subject, in each case, to the acceptability of the value obtained for the certified reference material.

### 8.2.3 Calculation of final result

The final result is the arithmetical mean of the acceptable analytical values of the test sample calculated to the fourth decimal place and rounded off to the second decimal place as follows:

1) When the figure in the third decimal place is less than 5, it is discarded and the figure in the second decimal place is kept unchanged.

2) When the figure in the third decimal place is 5 and there are figures other than 0 in the fourth decimal place, or when the figure in the third decimal place is greater than 5, the figure in the second decimal place is increased by one.

3) When the figure in the third decimal place is 5 and there are no figures other than 0 in the fourth decimal place, the 5 is discarded and the figure in the second decimal place is kept unchanged if it is 0, 2, 4, 6 or 8, and increased by one if it is 1, 3, 5, 7 or 9.

## 9 Test report

The test report shall include the following:

- details necessary for the identification of the sample;
- reference to this International Standard;
- results of the analysis;
- reference number of the results;
- any characteristics noticed during the determination and any operations not specified in the International Standard which may have had an influence on the results.

Table — Precision data for aluminium determinations

| Sample   | Ore type  | Mean Al <sub>2</sub> O <sub>3</sub> content<br>% (m/m) | Components of standard deviation |                | Reproducibility index<br>2s |
|----------|---|--|----------------------------------|----------------|-----------------------------|
|          |   |  | s <sub>w</sub>                   | s <sub>b</sub> |                             |
| MT/12/5  | Gibbsite (19,5 % Fe <sub>2</sub> O <sub>3</sub> )                             | 46,62  | 0,131                            | 0,212          | 0,50                        |
| MT/12/12 | Diaspore (21,7 % Fe <sub>2</sub> O <sub>3</sub> ,<br>4,0 % SiO <sub>2</sub> ) | 52,48  | 0,137                            | 0,306          | 0,67                        |
| MT/12/4  | Boehmite/gibbsite blend   | 53,39  | 0,120                            | 0,126          | 0,35                        |
| NBS 696* | Gibbsite  | 54,48  | 0,051                            | 0,149          | 0,31                        |

\* Certificate value 54,5 %.

where

s<sub>w</sub> is the within-laboratory standard deviation;

s<sub>b</sub> is the between-laboratories standard deviation.