



Standard Practice for Calculation of Supersaturation of Barium Sulfate, Strontium Sulfate, and Calcium Sulfate Dihydrate (Gypsum) in Brackish Water, Seawater, and Brines¹

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1. Scope

1.1 This practice covers the calculation of supersaturation of barium sulfate, strontium sulfate, and calcium sulfate dihydrate (gypsum) in brackish water, seawater, and brines in which barium, strontium, and calcium ions either coexist or exist individually in solution in the presence of sulfate ions.

1.2 This practice is not applicable for calculating calcium sulfate dihydrate supersaturation if the temperatures of saline waters under investigation exceed 95°C. At temperatures above 95°C, hemianhydrate and anhydrite would be major insoluble forms.

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 ASTM Standards:²

- D511 Test Methods for Calcium and Magnesium In Water
- D512 Test Methods for Chloride Ion In Water
- D513 Test Methods for Total and Dissolved Carbon Dioxide in Water
- D516 Test Method for Sulfate Ion in Water
- D1129 Terminology Relating to Water
- D3352 Test Method for Strontium Ion in Brackish Water, Seawater, and Brines

¹ This practice is under the jurisdiction of ASTM Committee D19 on Water and is the direct responsibility of Subcommittee D19.05 on Inorganic Constituents in Water.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

- D3370 Practices for Sampling Water from Closed Conduits
- D3561 Test Method for Lithium, Potassium, and Sodium Ions in Brackish Water, Seawater, and Brines by Atomic Absorption Spectrophotometry
- D3651 Test Method for Barium in Brackish Water, Seawater, and Brines
- D3986 Test Method for Barium in Brines, Seawater, and Brackish Water by Direct-Current Argon Plasma Atomic Emission Spectroscopy

3. Terminology

3.1 *Definitions*—For definitions of terms used in this practice, refer to Terminology D1129.

4. Significance and Use

4.1 This practice covers the mathematical calculation of the supersaturation of three principal sulfate scaling compounds found in industrial operations. Application of this standard practice to the prediction of scale formation in a given system, however, requires experience. The calculations tell the user if a water, or mixture of waters, is in a scaling mode. Whether or not scale will in fact form, how quickly it will form, where it will form, in what quantities, and what composition are subject to factors beyond the scope of this practice. However, based on how supersaturated a given water or mixture of waters is, an objective evaluation of the relative likelihood of scale formation can be made.

NOTE 1—There are several personal computer (PC) type programs that are both available commercially and publicly that will perform these calculations.

5. Procedure

5.1 Collect water samples for compositional analysis in accordance with Practices D3370.

5.2 Determine the calcium and magnesium concentrations in accordance with Test Methods D511.

5.3 Determine the barium concentration in accordance with Test Methods D3651 or D3986.

5.4 Determine the strontium concentration in accordance with Test Method D3352.

5.5 Determine sodium and potassium concentrations in accordance with Test Method **D3561**.

5.6 Determine sulfate ion concentration in accordance with Test Method **D516**.

5.7 Determine chloride ion concentration in accordance with Test Methods **D512**.

5.8 Determine carbonate and bicarbonate ion concentrations in accordance with Test Methods **D513**.

5.9 Determine the concentrations of all other major inorganic constituents that may be present in the water under investigation in accordance with appropriate test methods in *Annual Book of ASTM Standards*, Vols 11.01 and 11.02.

5.10 Determine temperature and pressure of the water system under investigation.

6. Calculation of Ionic Strength

6.1 Calculate the ionic strength of the water under investigation as follows:

$$\mu = \frac{1}{2} \sum C_i Z_i^2 \quad (1)$$

where:

μ = ionic strength,
 C_i = molal concentration of each ion in solution, and
 Z_i = charge number of ion, i.

7. Calculation of Barium Sulfate Supersaturation (Refer to Appendix X1)

7.1 Calculate barium sulfate solubility in the water under investigation, using the equation as follows:

$$S = (\sqrt{X^2 + 4K} - X)/2 \quad (2)$$

where:

S = solubility, moles of solute per kilogram of water corrected for the common ion effect,
 K = solubility product constant (molal) at the ionic strength, temperature and pressure of the water under investigation. For BaSO_4 refer to **Appendix X2**, and
 X = molal excess of soluble common ion.

7.2 Calculate the amount of barium sulfate, moles per kilogram of water, in the sample based on the lesser of the barium or sulfate ion concentration.

7.3 If the amount of BaSO_4 in the sample (7.2) is less than its calculated solubility (7.1), the water in question is undersaturated with respect to BaSO_4 . If the amount of BaSO_4 present is greater than its solubility, the water is supersaturated with respect to BaSO_4 . Calculate the amount of supersaturation as the difference between the two values:

$$\text{supersaturation} = \text{concentration} - \text{solubility} \quad (3)$$

NOTE 2—Supersaturation may also be calculated directly from the equation (1).³

$$([\text{Ba}^{++}] - y)([\text{SO}_4^{=}] - y) = K \quad (4)$$

where:

Ba^{2+} = concentration of barium, molal,
 SO_4^{2-} = concentration of sulfate, molal,
 y = excess (supersaturation) of BaSO_4 , molal, and
 K = solubility product constant (molal) of BaSO_4 at test conditions.

The value X may then be determined from the quadratic equation (see **Appendix X1**):

$$X = \frac{-B \pm \sqrt{B^2 - 4AC}}{2A}$$

Report BaSO_4 supersaturation in molal terms of the weight of BaSO_4 per volume of water, mg/L.

BaSO_4 supersaturation, mg/L

$$= \text{BaSO}_4, (\text{molal}^2) \times 10^3 \times 233 \times \left(\frac{1000 \times D}{\frac{TDS}{1000} + 1000} \right)$$

where:

D = sample density.

8. Calculation of Strontium Sulfate Supersaturation (Refer to Appendix X1)

8.1 Calculate strontium sulfate solubility using the same steps described for BaSO_4 (Section 7), but substituting the appropriate values for SrSO_4 in **Eq 2** (refer to **Appendix X3** or **Appendix X4**).

NOTE 3—If barium sulfate supersaturation exists, the amount of sulfate available for strontium sulfate will be less by the amount of sulfate equivalent to the calculated BaSO_4 supersaturation.

NOTE 4—If carbonate ions are present, strontium carbonate may precipitate. The amount of strontium may then be corrected by that required for strontium carbonate precipitation prior to the calculation of SrSO_4 solubility (2). Practically speaking, however, due to the extremely low solubility of SrCO_3 , this correction may usually be omitted.

8.2 Calculate the amount of strontium sulfate moles per kilogram water in the sample based on the lesser of the strontium or remaining sulfate ion concentration.

8.3 If the amount of SrSO_4 in the sample (8.2) is less than its calculated solubility (8.1), the water in question is undersaturated with respect to SrSO_4 . If the amount of SrSO_4 present is greater than its solubility, the water is supersaturated with respect to SrSO_4 . Calculate the amount of supersaturation, moles per kilogram water by difference (**Eq 3**), or by substituting appropriate data in **Eq 4** (**Note 2**).

8.3.1 Report SrSO_4 supersaturation in terms of the weight of SrSO_4 per volume of water as follows:

SrSO_4 supersaturation mg/L

$$= \text{SrSO}_4, (\text{molal}) \times 10^3 \times 184 \times \left(\frac{1000 \times D}{\frac{TDS}{1000} + 1000} \right)$$

9. Calculation of Calcium Sulfate Supersaturation (Refer to Appendix X1)

9.1 Calculate calcium sulfate solubility using the same steps described for BaSO_4 (Section 7), but substituting the appropriate values for CaSO_4 in **Eq 2** (refer to **Appendix X5**).

³ The boldfaced numbers in parentheses refer to a list of references at the end of this standard.

9.2 Calculate the amount of calcium sulfate moles per kilogram in the sample based on the lesser of the calcium or remaining sulfate ion.

9.3 If the amount of CaSO₄ in the sample (9.2) is less than its calculated solubility (9.1), the water in question is undersaturated with respect to CaSO₄. If the amount of CaSO₄ present is greater than its solubility, the water is supersaturated with respect to CaSO₄. Calculate the amount of supersaturation moles per kilogram by difference (Eq 3) or by substituting appropriate data in Eq 4 (Note 2).

9.3.1 Report CaSO₄ supersaturation in terms of the weight of CaSO₄·2H₂O (gypsum) per volume of water after converting moles per data obtained above to mg/L as follows:

$$\begin{aligned} & \text{CaSO}_4 \cdot 2\text{H}_2\text{O supersaturation, mg/L} \\ & = \text{CaSO}_4 \cdot 2\text{H}_2\text{O}_2, \text{ moles/kg} \times 172.17 \times 10^3 \times D \end{aligned}$$

10. Keywords

10.1 barium sulfate; brines; calcium sulfate dihydrate; strontium sulfate

APPENDIXES

(Nonmandatory Information)

X1. SAMPLE CALCULATION OF BaSO₄ SUPERSATURATION AT 95°C

Analysis of Water				Ionic Strength	
Component Ions	mg/L	moles per litre ^A	molal ^A Concentration	Z ²	μ = 1/2 ∑ _i Z _i ² (Section 6)
Na	27 120	1.180	1.214	1	1.214
Ca	10 890	0.272	0.280	4	1.120
Mg	1679	0.69	0.071	4	0.284
Ba	6.4	0.000044	4.52 × 10 ⁻⁵	4	>0.001
Sr	444	0.00506	521.42 × 10 ⁻⁵	4	0.021
Cl	64 870	1.830	1.883	1	1.883
SO ₄	1210	0.012596	1296.14 × 10 ⁻⁵	4	0.052
HCO ₃	317	0.005	0.005	1	0.005
TDS = 106 536 Density = 1.078 g/ml				Total ionic strength = 2.29 K _{BaSO₄} at 95° (Appendix X1) = 83.22 × 10 ⁻⁹	

A

$$\begin{aligned} \text{Convert moles/L to molal} &= \text{moles/L} \times \frac{1000}{(\text{Sp gr} \times 1000) - \frac{\text{TDS}}{1000}} \\ &= \text{moles/L} \times \frac{1000}{1078 - 106.5} \\ &= \text{moles/L} \times 1.029 \end{aligned}$$

X1.1 BaSO₄ Solubility (Refer to 7.1):

$$S = \left(\sqrt{X^2 + 4K} - X \right) / 2$$

where:

X = molal excess of common ion (in this case SO₄),

$$\begin{aligned} X &= (1296.14 \times 10^{-5}) - (4.52 \times 10^{-5}) \\ &= 1291.62 \times 10^{-5} \end{aligned}$$

$$4K = 4(83.22 \times 10^{-9}) = 332.88 \times 10^{-9}, \text{ or } 3328.8 \times 10^{-10}$$

$$S = \left[\sqrt{(1291.62 \times 10^{-5})^2 + (3328.8 \times 10^{-10})} - (1291.62 \times 10^{-5}) \right] / 2$$

$$\text{Solubility } S = 0.644 \times 10^{-5} \text{ molal}$$

X1.2 BaSO₄ Present (Refer to 7.2):

X1.2.1 Ba present = 4.52 × 10⁻⁵ molal

X1.2.2 SO₄ present = 1296.14 × 10⁻⁵ molal

X1.2.3 Based on lower value (Ba), BaSO₄ present = 4.52 × 10⁻⁵ molal

X1.3 Amount of BaSO₄ Supersaturation (Refer to 7.3):

X1.3.1 BaSO₄ present based on Ba²⁺ = 4.52 × 10⁻⁵ molal

X1.3.2 Calculated BaSO₄ solubility, S = 0.64 × 10⁻⁵ molal

X1.3.3 BaSO₄ excess; that is, supersaturation = 3.88 × 10⁻⁵ molal; or 8.8 mg/L of sample

X1.4 Useful Information:

	Mol Weight	Equivalent Weight	Gravimetric Conversion Factors
Ba	137.33	68.66	Ba × 1.6995 = BaSO ₄
Ca	40.08	20.04	Ca × 3.3967 = CaSO ₄
Sr	87.62	43.81	Sr × 2.0963 = SrSO ₄
SO ₄	96.06	48.03	
BaSO ₄	233.39	116.70	SO ₄ × 2.4296 = BaSO ₄
CaSO ₄	136.14	68.07	SO ₄ × 1.4172 = CaSO ₄
CaSO ₄ ·2H ₂ O	172.14	86.07	SO ₄ × 1.9121 = SrSO ₄
SrSO ₄	183.68	91.84	

X1.5 The amount of supersaturation (excess BaSO₄) may also be calculated directly using the expression (Eq 4):

$$([\text{Ba}^{++}] - X) ([\text{SO}_4] - X) = K_{\text{BaSO}_4}$$

X1.5.1 Using the molal values from the water analysis above this becomes:

$$([4.52 \times 10^{-5}] - X) ([1296.14 \times 10^{-5}] - X) = 832.2 \times 10^{-10}$$

Multiplying: $(5858.55 \times 10^{-10}) - (1300.66 \times 10^{-5})$

$$X + X^2 = 832.2 \times 10^{-10}$$

X1.5.2 Substituting the above coefficients of X in the quadratic equation:

$$X = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

Combining: $X^2 - (1300.66 \times 10^{-5}) X + 5026.35 \times 10^{-10} = 0$

and solving, $X = 3.88 \times 10^{-5}$ molal; or 8.8 mg/L of sample.

X2. SOLUBILITY DATA FOR BaSO₄·NaCl·H₂O SYSTEMS (3)

Solution Ionic Strength, μ	Solubility Product Constant, K (Molal)					
	25°C	35°C	50°C	65°C	80°C	95°C
0.1	1.54×10^{-9}	2.00×10^{-9}	2.70×10^{-9}	3.34×10^{-9}	3.76×10^{-9}	3.97×10^{-9}
0.2	2.70	3.36	4.76	5.93	7.06	7.74
0.4	4.49	5.63	7.92	10.61	13.69	16.13
0.6	6.08	7.74	11.03	15.38	20.45	24.97
0.8	7.74	9.60	13.69	20.16	26.57	33.49
1.0	9.22	11.24	16.38	24.02	32.76	42.02
1.5	12.54	15.38	22.20	32.40	44.94	62.00
2.0	15.63	19.04	27.23	39.60	56.17	78.96
2.5	18.23	21.90	31.33	44.94	63.50	93.64
3.0	20.74	24.65	34.97	49.73	70.23	107.57
3.5	23.41	27.56	38.81	53.82	76.73	120.41
4.0	25.92	30.63	42.44	58.08	82.94	132.50
4.5	28.56	34.23	45.80	63.00	89.40	144.40

X3. SOLUBILITY PRODUCT DATA FOR SrSO₄·2NaCl·H₂O SYSTEMS (4)

Solution Ionic ^A Strength, μ	Solubility Product Constant, K (Molal)	
	40°C (104°F)	71°C (160°F)
0.1	0.250×10^{-5}	0.160×10^{-5}
0.2	0.390	0.250
0.3	0.505	0.345
0.4	0.617	0.440
0.5	0.723	0.518
0.75	1.02	0.785
1.0	1.26	1.04
1.25	1.48	1.25
1.5	1.68	1.41
1.75	1.86	1.57
2.0	2.00	1.68
2.25	2.09	1.76
2.5	2.14	1.81
2.75	2.16	1.84
3.0	2.17	1.86
3.25	2.19	1.87
3.50	2.20	1.88

^A The above table may be used to interpolate the solubility product (K) for SrSO₄ in brines at 0 psig. The interpolated values can be substituted in Eq 2 (Section 7) for estimating the solubility (S) of SrSO₄. For more precise K values at temperatures up to 300°F (149°C) and pressures up to 3000 psig add SI unit, refer to Appendix X4.

X4. EQUATION FOR CALCULATING SrSO₄ SOLUBILITY (5)

X4.1 Experimental SrSO₄ solubility data have been reduced to the following regression equation for calculating the solubility product constant (*K*) at various solution ionic strengths over a temperature range of 100 to 300°F (38 to 149°C) and pressures up to 3000 psig. The equation is adaptable to computer calculation which can then substitute the value for *K* in Eq 2 (Section 7) for computing the solubility of SrSO₄ at desired conditions.

$$\text{Log } K_{\text{SrSO}_4} = X/R$$

where:

$$X = 1/T,$$

$$R = A + BX + C\mu^{1/2} + D\mu + EZ^2 + FXZ + G\mu^{1/2}Z,$$

Z = pressure (psig),

μ = solution ionic strength,

T = temperature, °K.

X4.1.1 Coefficients of the above equation for *R* are as follows:

$$A = 0.266948 \times 10^{-3}$$

$$B = -244.828 \times 10^{-3}$$

$$C = -0.191065 \times 10^{-3}$$

$$D = 53.543 \times 10^{-6}$$

$$E = -1.383 \times 10^{-12}$$

$$F = 1.103323 \times 10^{-9}$$


$$G = -0.509 \times 10^{-9}$$

X5. SOLUBILITY PRODUCT DATA FOR CaSO₄·2-NaCl·H₂O SYSTEMS (6)

Solution Ionic Strength, μ	Solubility Product Constant, <i>K</i> (Molal)			
	10°C	35°C	50°C	80°C
0	1.02×10^{-4}	1.27×10^{-4}	1.25×10^{-4}	0.89×10^{-4}
0.1	3.04	3.29	3.31	2.82
0.2	4.99	5.23	5.28	4.67
0.3	6.87	7.11	7.17	6.44
0.4	8.68	8.91	8.96	8.13
0.5	10.41	10.64	10.68	9.75
0.6	12.07	12.30	12.30	11.30
0.7	13.65	13.88	13.85	12.78
0.8	15.16	15.39	15.32	14.18
0.9	16.60	16.83	16.71	15.52
1.0	17.96	18.20	18.02	16.79
1.25	21.05	21.29	20.96	19.70
1.5	23.69	23.93	23.46	22.22
1.75	25.90	26.12	25.52	24.39
2.0	26.67	27.88	27.18	26.22
2.25	29.03	29.22	28.47	27.73
2.5	30.00	30.15	29.40	28.92
2.75	30.60	30.71	30.01	29.81
3.0	30.84	30.90	30.32	30.42
3.25	30.77	30.77	30.36	30.73
3.5	30.39	30.34	30.15	30.76
3.75	29.76	29.66	29.73	30.51
4.0	28.90	28.75	29.13	29.97
4.25	27.85	27.66	28.37	29.14
4.5	26.65	26.43	27.49	28.02
4.75	25.34	25.13	26.52	26.58
5.0	23.98	23.80	25.48	24.83

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